



How do diamonds grow in metal melt together with silicate minerals? An experimental study of diamond morphology

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Abstract. The origin and evolution of metal melts in the Earth's mantle and their role in the formation of diamond are the subject of active discussion. It is widely accepted that portions of metal melts in the form of pockets can be a suitable medium for diamond growth. This raises questions about the role of silicate minerals that form the walls of these pockets and are present in the volume of the metal melt during the growth of diamonds. The aim of the present work was to study the crystallization of diamond in a complex heterogeneous system: metal–melt–basalt–carbon. The experiments were performed using a multianvil high-pressure apparatus of split-sphere type (BARS) at a pressure of 5.5 GPa and a temperature of 1500 °C. The results demonstrated crystallization of diamond in metal melt together with garnet and clinopyroxene, whose chemical compositions are similar to those of eclogitic inclusions in natural diamond. We show that the presence of silicates in the crystallization medium does not reduce the chemical ability of metal melts to catalyze the conversion of graphite into diamond, and, morphologically, diamond crystallizes mainly in the form of a cuboctahedron. When the content of the silicate material in the system exceeds 5 wt %, diamond forms parallel-growth aggregates, but 15 wt % of silicate phases block the crystallization chamber, preventing the penetration of metallic melt into them, thus interrupting the growth of diamond. We infer that the studied mechanism of diamond crystallization can occur at lower-mantle conditions but could also have taken place in the ancient continental mantle of the Earth, under reducing conditions that allowed the stability of Fe–Ni melts.

1 Introduction

Natural diamond crystals are the best containers for storing and transporting material from great depths. Mineral inclusions in diamond may thus in many cases be indicators of the diamond crystallization environment. In lithospheric diamonds, the genesis of which is associated with mantle depths of less than 250 km, inclusions of silicate minerals such as olivine, garnet, and pyroxene are common (Sobolev, 1977; Harris and Gurney, 1979; Gurney, 1989; Sobolev et al., 1973, 2004; Stachel and Harris, 2008). In most cases, these inclusions belong to the peridotite assemblage, although available data also indicate a significant role of inclusions of eclogitic assemblage (Gurney, 1984). Moreover, besides quantitative predominance, diamonds with peridotite inclusions are also the most ancient (Gurney et al., 2010). For super-deep, so-called sublithospheric diamonds, there is much less informa-

tion, although these diamonds are increasingly attracting the interest of researchers (Stachel, 2001; Stachel et al., 2005; Kaminsky et al., 2009; Harte, 2010; Thomson et al., 2014; Pearson et al., 2014; Zedgenizov et al., 2016; Smith and Wang, 2017; Nestola, 2017; Nestola et al., 2018; Smith et al., 2018; Anzolini et al., 2018, 2019).

Metallic inclusions are a specific type of inclusion in natural diamonds from kimberlites (Sobolev et al., 1981; Meyer and McCallum, 1986; Vinokurov et al., 1998; Garanin and Kudryavtseva, 1990; Bulanova and Zayakina, 1991; Bulanova, 1995; Gorshkov et al., 1997; Bulanova et al., 1998; Stachel et al., 1998; Win et al., 2001; Hayman et al., 2005; Bulanova et al., 2010; Kaminsky and Wirth, 2011; Kaminsky, 2012; Kagi et al., 2016). These inclusions were found in lithospheric diamonds. The interest in the study of metallic inclusions significantly increased after their finding in large

crystals of sublithospheric diamond, the so-called CLIPPIR diamonds (Smith et al., 2016), even if the relationship between metal melts and the genesis of natural diamond has been increasingly discussed (Smith and Kopylova, 2014; Smith et al., 2017). A characteristic feature of CLIPPIR diamonds is a variably light carbon isotopic composition, consistent with that of eclogitic diamonds and thought to result from the recycling of crustal carbon. The assumption that the carbon source for eclogitic diamonds is crustal material subducted to mantle depths seems quite possible, because the phenomenon of reduction of carbonate carbon to elemental carbon during subduction was proposed earlier (Malvoisin et al., 2012; Galvez et al., 2013) as well as demonstrated experimentally in the presence of metallic iron with the resulting graphite or diamond depending on pressure (Chepurov et al., 2011; Palyanov et al., 2013; Martirosyan et al., 2016).

The model of the diamond formation in the deep mantle is represented as the crystallization of diamonds in metal-melt pockets in host rocks (Smith et al., 2016). This raises questions about the role of silicate minerals in the process of crystallization of such diamonds – it is obvious that the evolving metal melts could interact in a certain way with the silicate rocks that enclose these metal pockets. Also, silicate minerals in the form of individual crystals and their aggregates were present in the volume of the metal melt during the dynamic evolution of metal pockets and directly during the growth of diamond crystals. Thus, many aspects still seem unclear about the crystallization medium of diamond, which is a complex heterogeneous system consisting of metal melts in which solid silicate phases are present. In this work, we studied synthetic diamond crystals and silicate minerals crystallized together in a metal-melt-basalt-carbon system. The main questions addressed were the transformation of the morphology of diamond crystals growing on a seed depending on the seed location in different zones of a reaction chamber, as well as the possibility of complete blocking of diamond growth through addition of various amounts of silicates to the system.

2 Experimental methods

The experiments were performed on a multianvil high-pressure apparatus of split-sphere type (BARS) at the Institute of Geology and Mineralogy of the Siberian Branch of the Russian Academy of Sciences in Novosibirsk according to state assignment 0330-2016-0012. As the high-pressure cells, containers made from a mixture based on ZrO_2 were used. The rectangular parallelepiped pressure cells were 20 mm × 20 mm × 23 mm in size and had truncated edges and vertices. The tube-shaped heater was made of graphite. As electrical contacts in the cell, parts made of molybdenum were used. The pressure in the cell was calibrated at room temperature with the reference substances Bi and PbSe (Turkin, 2003, 2004). Temperature was mea-

sured with a PtRh₃₀–PtRh₆ thermocouple without a pressure correction. The temperature correction for pressure was determined by the melting curve of pure Ag and Au (Tonkov and Ponyatovsky, 2004). The measurement error is +0.2 GPa and +25 °C. Heating of the samples was performed after pressurization, and after the experiment the samples were cooled by quenching. Further experimental details with photographs are available in Chepurov et al. (2010), Zhimulev et al. (2012), and Tomilenko et al. (2015).

An experimental study of diamond crystallization in model systems was performed using several configurations of high-pressure cell assembly (Fig. 1a–d). The carbon source (graphite) was placed in the central, higher-temperature zone, and the metal–silicate charge and the diamond seed crystals were placed above or below it. The choice of the configurations made it possible to study the crystallization of diamond on the seeds, located in different parts of the reaction volume, and to evaluate the effect of the initial arrangement of silicate additives on their capture by the growing diamond crystal and on the change of diamond morphology. A synthetic diamond mesh with a crystal size of 0.3 mm was chosen for the seed crystals. The metal component of the sample was a mixture of carbonyl iron and nickel in a mass ratio of 36 : 64. The silicate additive to the metal mixture was a fine-grained (5 μm) powder of natural alkaline basalt glass (wt %): SiO₂, 46.60; TiO₂, 2.24; Al₂O₃, 15.27; Fe₂O₃, 4.82; FeO, 6.38; MnO, 0.15; CaO, 6.68; MgO, 4.25; NaO, 6.14; K₂O, 4.70; P₂O₅, 2.34; LOI, 0.1; total, 99.67. According to our planning of the experiment, at the chosen *P–T* parameters the basalt forms an eclogitic mineral assemblage.

The experiments were carried out at a pressure of 5.5 GPa and a temperature of 1500 °C, i.e., within the diamond stability field (Day, 2012), for a duration of 72 h. Such a duration was chosen in order to grow diamond single crystals weighing up to 0.5 carat, which are evidence of the crystallization of diamond macrocrystals. Basalt was added to each experiment in amounts of 1 wt %–15 wt %. Samples after experiments were extracted from the cell, and the metallic component was dissolved in acids.

The surface morphology of the obtained diamond crystals was investigated using optical (MC2-Zoom, Olympus BX35) and scanning electron microscopy (1540 XB Cross-beam, Carl Zeiss), as well as confocal laser scanning microscopy (LSM 800MAT, Carl Zeiss) in CKP “Nanostruktury” (Novosibirsk, Russia). The chemical composition of the mineral phases was determined using a MIRA LMU scanning electron microscope and a JEOL JXA-8100 micro-analyzer. The measurements were carried out at an accelerating voltage of 15–20 kV, a probe current of 20–50 nA, and a counting time of 10–60 s. To determine the composition of homogeneous phases, a probe with a minimum spot diameter of about 3 μm was used. To determine the average composition of the phase mixture, a defocused probe with a diameter of 20 to 100 μm was used (depending on the grain size in the mixture).

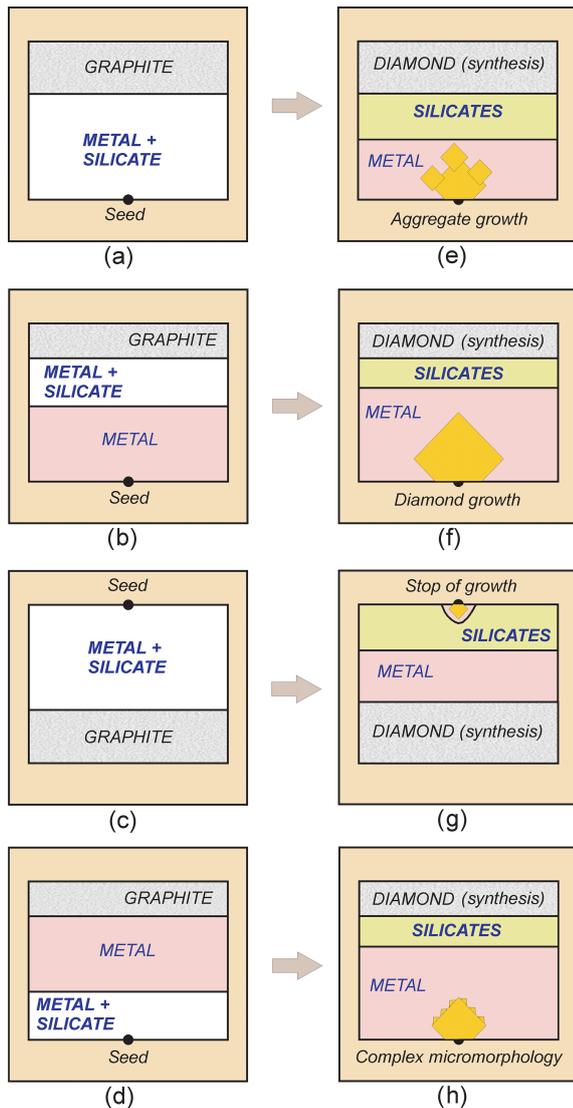


Figure 1. Schematic vertical cross sections of high-pressure sample assembly configurations used in experiments for diamond growth together with silicate minerals (a–d), and sketch drawing of the respective run products (e–h). (a) Basalt is evenly mixed with metal powder, and the diamond seed is located at the bottom of the chamber; (b) similar configuration but the lower part of initial metal does not contain basalt; (c) basalt is evenly mixed with metal powder, and the diamond seed is located at the top of the chamber; (d) the diamond seed is located at the bottom of the chamber and is covered with a thin layer of a mixture of metal and basalt – the rest of the sample is a basalt-free Fe–Ni metal. Note that after experiments newly formed silicate phases were found mainly in the upper part of the chamber. Growth of diamond on a seed crystal occurred in the form of monocrystals (f), crystals of complex micromorphology with growth pyramids present on (111) faces (h), or as parallel-growth aggregates (e). Only a little diamond growth was observed when the seed crystals were placed at the top of the chamber (g). For further explanation of run products the reader is referred to the Results section and to Table 1.

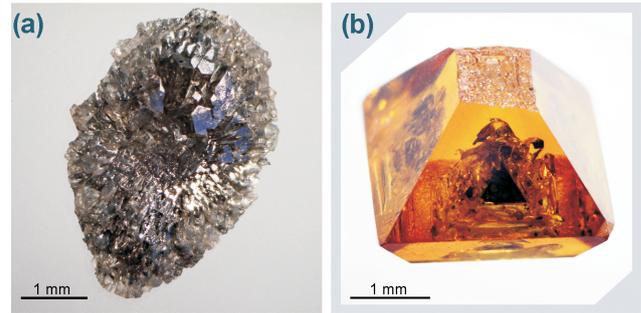


Figure 2. (a) General view of the carbon source residue after run 2–41, which is a crystalline diamond aggregate; (b) diamond crystal weighing 0.5 carat from run 2–41, grown in the Fe–Ni–C system with 1 wt% silicates added to the system (assembly configuration shown in Fig. 1a). The top of the crystal was in contact with the chamber wall and does not have faces. The crystal contains well-shaped inclusions of metal.

3 Results

The samples recovered at the end of experiments contained a quenched metal alloy, synthetic diamond, and synthetic silicates (Table 1). After dissolving the metal alloy, it was found that graphite at the source (Fig. 1e–h) completely turned into diamond (synthesis) and consisted of a finely crystalline diamond aggregate (Fig. 2a), with individual diamond grains 0.1–0.2 mm in size. The carbon source was partially dissolved and part of the carbon was transported to the growing seed crystal, which is a typical case for high-pressure, high-temperature (HPHT) processes.

Growth of a diamond on the seed crystals was observed in most experiments, primarily when the seeds were placed at the lower part of the assembly, and the added silicates were initially located at the upper part of the chamber (Fig. 1b). In this case (Fig. 1f) we observed the formation of diamond single crystals typical of diamonds grown in the Fe–Ni–C system. A more complicated situation was observed in experiments when the seed crystals were in contact with metal that initially contained a silicate component (Fig. 1a, d). In experiments with a silicate additive content of less than 5 wt%, diamonds grown on the seeds were high-quality crystals of octahedral habit with the additional faces {110}, {311}, and {100} (Fig. 2b). The crystal faces were free of surface sculptures, and in rare cases dendritic and network sculptures were observed. The color of the diamond was yellow. When the content of silicates in the system was 5 wt%–10 wt%, a number of specific features of the micromorphology of diamond crystals were observed. Growth pyramids with triangular forms, which have the same orientation as the triangle of the octahedral face, were found on diamond octahedral faces (Fig. 3a). The pyramids are faceted with stepped surfaces, most likely a trigon-trioctahedron. Measuring the profile of the pyramids by confocal laser scanning microscopy revealed that the vertical size of most of the pyramids was

Table 1. Experimental conditions and run products.

Run no.	<i>P</i> , GPa	<i>T</i> , °C	Duration, hours	Starting sample	Assembly cofig.*	Basalt content**	Run products	Morphology of Dia _{Gr}
2–41	5.5	1500	72	Fe–Ni, C, Bas	a	1	Fe–Ni–C, Dia _{Syn} , Dia _{Gr} , Gar, Cpx	Crys
2–42	5.5	1500	72	Fe–Ni, C, Bas	a	3	Fe–Ni–C, Dia _{Syn} , Dia _{Gr} , Gar, Cpx	Crys
2–43	5.5	1500	72	Fe–Ni, C, Bas	a	5	Fe–Ni–C, Dia _{Syn} , Dia _{Gr} , Gar, Cpx	Crys + GPyr
2–44	5.5	1500	72	Fe–Ni, C, Bas	a	10	Fe–Ni–C, Dia _{Syn} , Dia _{Gr} , Gar, Cpx	Agg
2–45	5.5	1500	72	Fe–Ni, C, Bas	a	15	Fe–Ni–C, Dia _{Syn} , Dia _{Gr} , Gar, Cpx	Agg
2–46	5.5	1500	72	Fe–Ni, C, Bas	b	10	Fe–Ni–C, Dia _{Syn} , Dia _{Gr} , Gar, Cpx	Crys
2–47	5.5	1500	72	Fe–Ni, C, Bas	b	15	Fe–Ni–C, Dia _{Syn} , Dia _{Gr} , Gar, Cpx	Crys + GPyr
2–48	5.5	1500	72	Fe–Ni, C, Bas	c	5	Fe–Ni–C, Dia _{Syn} , Dia _{Gr} , Gar, Cpx	Agg
2–49	5.5	1500	72	Fe–Ni, C, Bas	c	10	Fe–Ni–C, Dia _{Syn} , Dia _{Gr} , Gar, Cpx	Agg
2–50	5.5	1500	72	Fe–Ni, C, Bas	c	15	Fe–Ni–C, Dia _{Syn} , Gar, Cpx	No growth
2–51	5.5	1500	72	Fe–Ni, C, Bas	d	5	Fe–Ni–C, Dia _{Syn} , Dia _{Gr} , Gar, Cpx	Crys + GPyr
2–52	5.5	1500	72	Fe–Ni, C, Bas	d	10	Fe–Ni–C, Dia _{Syn} , Dia _{Gr} , Gar, Cpx	Agg

Fe–Ni – initial Fe_{0.36}Ni_{0.64} powder; C – initial graphite source; Bas – basalt powder; Fe–Ni–C – metal melt; Dia_{Syn} – diamond synthesis in source; Dia_{Gr} – diamond growth on seed; Gar – garnet; Cpx – clinopyroxene; Crys – diamond monocrystal; Crys + GPyr – diamond monocrystal with growth pyramids; Agg – parallel-growth aggregate of diamonds. * Assembly configuration according to Fig. 1. ** Content of basalt in metal–basalt mixture only, according to assembly configuration shown in Fig. 1 (wt %).

in the range 1–3 μm (Fig. 3b, c). Many such pyramids were often observed on the octahedral faces (Fig. 3d), and also numerous caverns were present which had well-cut contours (Fig. 3e). This effect was characteristic to varying degrees for all diamonds grown on the seed crystals from experiments with silicate content greater than 5 wt %.

In experiments with the seed crystals located at the bottom of the chamber (Fig. 1a), and when the content of the silicate component in the metal reached 10 wt %–15 wt %, diamond crystals grown on the seeds had a very complex surface relief and in most cases formed a parallel-growth aggregate of numerous individuals (Fig. 4). On the surface of diamonds from these experiments, the presence of a silicate material in the form of fine-grained, light-colored aggregates was observed, which filled the interstices between individuals (Fig. 5a). An important observation was noted for experiments with the seed crystals placed at the top of the chamber (Fig. 1c): in most cases, with the addition of silicates to the system, it was not possible to grow a representative diamond crystal on the seeds, but often with only a little growth observed (Fig. 1g), and with the addition of 15 wt % of silicate material, no growth of diamond on the seeds occurred.

Grown diamonds often contained dark faceted inclusions typical of diamonds produced in the Fe–Ni–C system (Fig. 2b). Another type of inclusions was identified in diamonds which is completely nontypical for crystals grown in metal–carbon systems, namely aggregates of numerous fine-grained phases (Fig. 5b, c). These inclusions have light-gray color. The typical size of individual aggregates is 100–200 μm (Fig. 5c). These aggregates are most often elongated along the octahedron growth sector of diamond, passing through almost the entire crystal (Fig. 5b). It has often been observed that these inclusions reached the surface of diamond – in this place a cavernous area was observed, filled

with light-gray material (Fig. 5d), similar to what is found on the surface of diamond between individuals (Fig. 5a).

The silicate phases after the experiments were found mainly in the upper part of the sample. It was a fine-grained aggregate which consisted of garnet and clinopyroxene (Fig. 6). Small drops of quenched metal were observed between grains of the silicate minerals (Fig. 6b). Representative chemical analyses are shown in Tables 2 and 3. The synthesized silicates are characterized by stable compositions, and we found no significant difference in the chemical composition of minerals crystallized in experiments with different amounts of basalt added to the system. Clinopyroxene is characterized by a high content of sodium (up to 9.82 wt % Na₂O) and potassium (0.14 wt %–0.31 wt % K₂O). An impurity of titanium (0.91 wt %–1.42 wt % TiO₂) is characteristic in pyroxene. The iron content in pyroxene is low (5.37 wt %–7.48 wt % FeO). Garnet is characterized by a stable content of CaO and FeO, 11.01 wt %–14.75 wt % and 17.57 wt %–23.27 wt %, respectively. Garnet has low chromium content (0.09 wt %–0.37 wt % Cr₂O₃) but a rather high titanium content that reaches 2.12 wt % TiO₂. It should be noted that increasing the silicate component did not change the composition of the final silicate phases: with an increase of the basalt addition to the initial charge varying from 1 wt % to 15 wt %, the amount of silicates in the experimental products increased markedly, but their FeO content did not change.

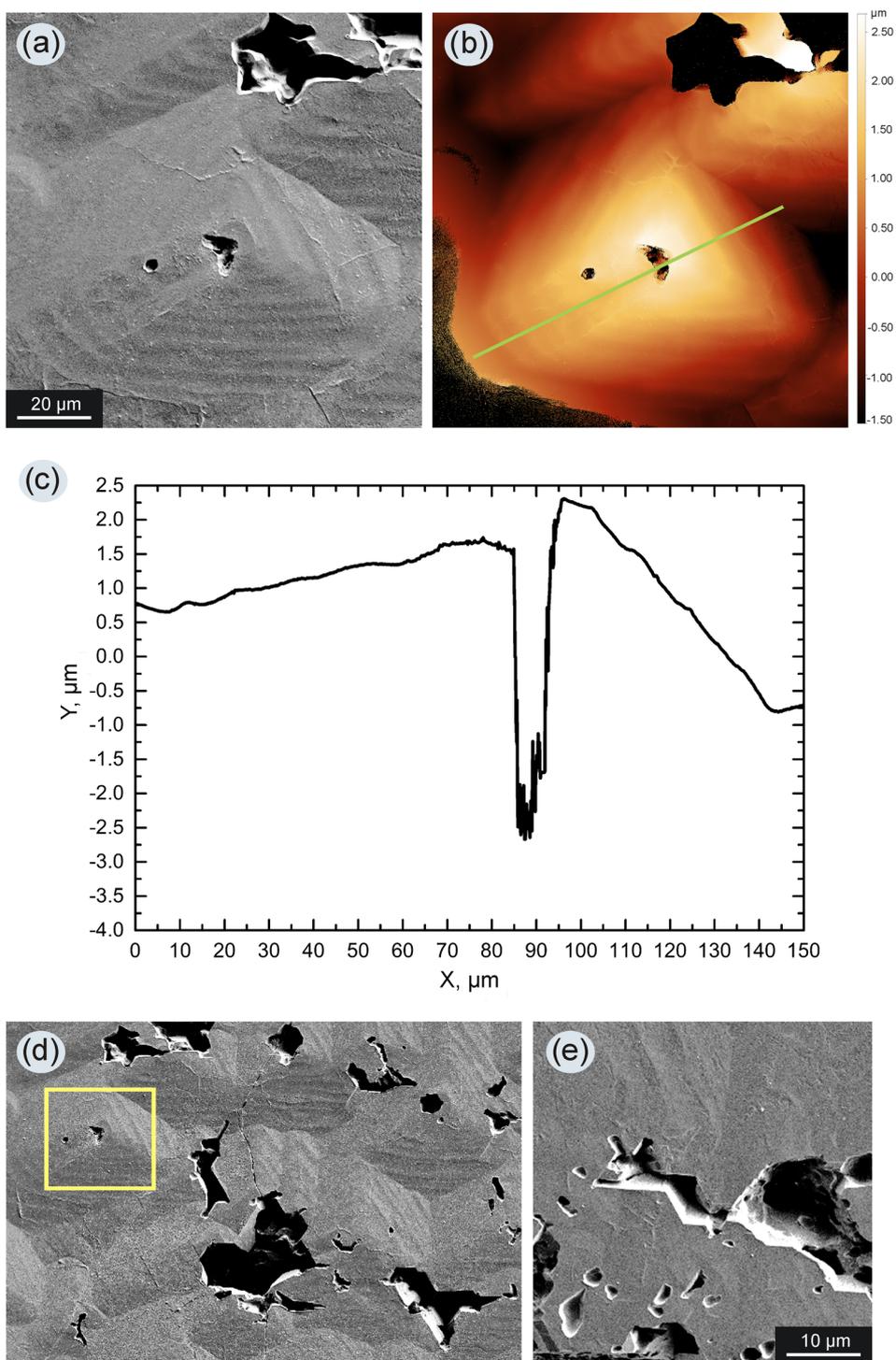


Figure 3. (a) SEM image of a growth pyramid on an octahedral face of diamond from run 2–51 with the addition of 5 wt % silicates (assembly configuration shown in Fig. 1d) – a cavern is present in the top of the pyramid; (b) relief of the pyramid shown in (a), measured by confocal laser scanning microscopy and the pyramid profile (c); (d) on (111) faces of the diamond, numerous pyramids are usually observed; the light-colored square indicates the pyramid that is shown in (a); (e) caverns on (111) diamond faces commonly have partially faceted edges.

Table 2. Representative chemical analyses of garnet from the experiments.

Sample	2-41/G1	2-42/G3	2-43/G1	2-44/G4	2-45/G2	2-46/G6	2-47/G3	2-48/G7	2-49/G5	2-50/G2
SiO ₂	39.02	39.18	39.17	39.34	39.28	39.24	39.09	39.28	39.04	39.03
TiO ₂	1.63	2.12	0.98	1.84	2.11	2.09	1.78	1.88	1.50	1.69
Al ₂ O ₃	20.01	20.05	20.83	19.37	19.09	19.49	19.76	19.86	19.57	20.55
Cr ₂ O ₃	0.37	0.25	0.35	0.29	0.18	0.09	0.16	0.33	0.29	0.12
FeO	20.94	17.57	19.15	19.54	20.66	19.66	22.25	18.89	23.27	18.12
MnO	0.35	0.29	0.34	0.30	0.35	0.30	0.29	0.31	0.31	0.30
MgO	5.21	5.66	6.22	4.45	5.33	5.26	4.85	5.83	4.76	5.84
CaO	12.04	14.51	12.83	14.75	12.86	13.55	11.71	13.32	11.01	14.05
Total	99.57	99.63	99.87	99.88	99.86	99.68	99.89	99.70	99.75	99.70
Si	3.025	3.010	3.006	3.044	3.041	3.033	3.034	3.023	3.044	2.998
Ti	0.095	0.122	0.057	0.107	0.123	0.121	0.104	0.109	0.088	0.098
Al	1.828	1.815	1.884	1.767	1.742	1.776	1.808	1.802	1.799	1.860
Cr	0.023	0.015	0.021	0.018	0.011	0.006	0.010	0.020	0.018	0.007
Fe	1.358	1.129	1.229	1.265	1.338	1.271	1.444	1.216	1.517	1.164
Mn	0.023	0.019	0.022	0.020	0.023	0.020	0.019	0.020	0.020	0.020
Mg	0.602	0.648	0.712	0.513	0.615	0.606	0.561	0.669	0.553	0.669
Ca	1.000	1.194	1.055	1.223	1.067	1.122	0.974	1.098	0.920	1.156
Total	7.954	7.953	7.985	7.956	7.960	7.955	7.953	7.957	7.960	7.971
100Fe#	69.28	63.52	63.33	71.13	68.50	67.71	72.02	64.51	73.28	63.51

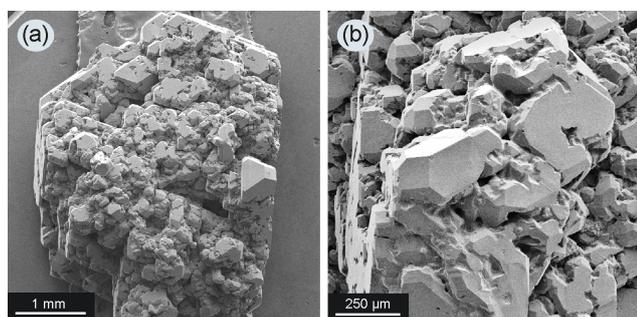


Figure 4. Diamond crystal grown in the Fe–Ni–C system with the addition of 10 wt % silicates (run 2–44, assembly configuration shown in Fig. 1a): (a) while maintaining the appearance of an octahedral crystal, the diamond grew as a parallel aggregate of numerous individuals; (b) enlarged fragment of the same crystal – each diamond subindividual has numerous faceted caverns on its faces.

4 Discussion

4.1 Morphology of diamond crystals grown on seed crystals in a metal melt together with silicate minerals

The experiments performed demonstrated the crystallization of diamond in the metal–carbon system together with silicate minerals. During the first minutes of the experiment, in the diamond *P–T* stability field, graphite in the source turned into diamond (synthesis), i.e., there was homogeneous nucleation (Giardini and Tydings, 1962). Next, the process of crystallization of diamond on a seed crystal began (heteroge-

neous nucleation), where the driving force is the temperature gradient in the pressure cell, in other words, the concentration of carbon in metal melt. In most cases, diamonds grown on the seeds have an octahedral habit with additional faces of a cube typical for crystals grown in metal–carbon systems. This is very important because it shows that, in metal melts in which various silicate components are present, morphologically typical octahedral crystals may grow.

Meanwhile, our study showed a number of specific macro- and micromorphological features characteristic of diamonds grown in the presence of silicates in the system. Diamond crystals have flat faces only in experiments with small additions of silicate components to the system, sometimes with dendritic and network sculptures which are well-known for diamonds synthesized in metal–carbon systems (Tolansky, 1962; Kanda et al., 1980; Field, 1992). Such sculptures on diamond faces are imprints of metal phases which crystallize as the metal melt cools.

The effect of silicate components on the morphology of diamond crystals starts to appear when their amount is 5 wt % in the crystallization medium. In the first stage, we observed growth of pyramids with triangular forms having the same orientation as the triangle of the octahedral face (Fig. 3). Even at this initial stage, the face of a diamond acquires a relief with an average height of more than 3 μm (Fig. 3c), whereas high-quality as-grown synthetic HPHT diamond crystals usually have a surface roughness of less than 5 nm (Chepurov et al., 2018a). Since several such pyramids formed on the faces of the crystal, after a certain time of growth, the face of a diamond crystal splits into several seg-



Figure 5. Diamond crystals with silicate inclusions: (a) part of a diamond crystal from a parallel-growth aggregate (run 2–44) – dark-colored arrows indicate the silicates in the form of fine-grained, light-colored aggregates that are intergrown with the diamond on its surface; silicates also fill the interstices between the diamond individuals; (b) diamond crystal of octahedral habit (run 2–43) – the white square shows a window where the chains of fine-grained silicate aggregates are visible inside the crystal; (c) the dark-colored arrow shows the inclusion of silicates of rounded shape consisting of a fine-grained aggregate; the inclusion is located close below the surface of the {111} diamond face (run 2–42); (d) silicate inclusion (shown by the light-colored arrow) that is not completely overgrown by diamond in run 2–42 and has a specific appearance on the surface in the form of cavities filled with light-colored fine-grained material (shown by dark arrow).

ments corresponding to individual pyramids. Thus, the formation of individual pyramids on the facets of crystals is associated with the occurrence of independent centers for new growth layers. Further growth of single crystals became impossible, and several subindividuals formed on the initial diamond are further developed into a parallel-growth aggregate (Fig. 4). Such diamonds were often observed in our experiments.

The results of the experiments were largely affected by the type of assembly we used. When the diamond seed crystal was placed at the bottom of the chamber, and the basalt powder was present only in the upper part (Fig. 1b), diamond monocrystals acquire a complicated morphology only when the diamond had reached the upper zone where silicates were present. When the seed diamond was located in the upper part of the reaction volume (Fig. 1c), diamond crystals with a significant amount of defects and the complicated morphology began to form even with the addition of small amounts

of silicates in the mixture. Due to differences in density and insignificant solubility in the metal melt, the silicates in the experiments float to the top of the chamber and shield the surface of growing diamond crystals. Diamond growth on the seeds was almost completely stopped at a content of 15 wt % of silicates in the system.

Comparing the crystals synthesized in experiments with natural diamonds, the following can be noted. It is known that the octahedral faces of natural diamonds commonly have a so-called polycentric structure of faces. It is believed that this sculpture is the result of the appearance of several growth centers on the face of a crystal (Bartoshinsky and Kvasnitsa, 1991). A similar phenomenon took place in our experiments; however, unlike the typical surface sculpture above composed of parallel triangular protrusions, separate growth pyramids with stepped side surfaces were present on diamonds grown in our experiments. Similar growth pyramids with triangular shapes were found on octahedral faces of natural diamond crystals that did not experience or had very little dissolution (Sunagava et al., 1984; Van Enckevort, 1992). Their formation is associated with groups of screw-type dislocations and is considered as evidence of the spiral-layer mechanism of their growth. It is likely that the capture of mechanical impurities by the crystal causes the appearance of stepped pyramids, and we observed, in the apex of each pyramid (Fig. 3a–c), the presence of a cavity with irregular contours, which is most likely the original location of the impurity on the diamond face. On the {111} faces of synthetic diamonds, similar accessories are also noted (Shigley et al., 1987). Thus, it was experimentally demonstrated that the most intensive formation of growth pyramids occurs when diamond grows from metal melts in which silicon is present (Kanda and Ohsawa, 1996). We observed a similar phenomenon when silicate components were added to the metal system.

Parallel-growth aggregates of diamond crystals observed in our experiments are of particular interest because research on diamonds with complex morphology is almost absent in the literature. It has been known for a long time that parallel-growth aggregates of crystals are commonly observed among diamonds from kimberlites (Fig. 7) (Orlov, 1963), but their genesis and crystallization conditions are still unclear. Our experiments suggest that in a complex heterogeneous system when diamond crystallizes from a metallic liquid, crystals with a very complex morphology can form, including parallel-growth aggregates of diamond similar to those found in nature.

4.2 Silicate phases crystallized together with diamond in a (Fe–Ni) metal–carbon melt

In the experiments performed, the newly formed silicate phases were found in the volume of the metal sample, on the surface of diamond crystals, and as inclusions in diamond. This means that when studying silicate phases from our ex-

Table 3. Representative chemical analyses of clinopyroxene from the experiments.

Sample	2–41/G1	2–42/G3	2–43/G1	2–44/G4	2–45/G2	2–46/G6	2–47/G3	2–48/G7	2–49/G5	2–50/G2
SiO ₂	54.18	54.04	54.39	54.61	54.12	54.27	53.85	54.13	54.20	54.48
TiO ₂	1.05	1.11	0.94	1.06	1.04	1.16	1.30	1.42	0.96	0.91
Al ₂ O ₃	15.58	13.67	14.47	14.79	15.60	15.31	14.37	16.28	15.21	14.89
Cr ₂ O ₃	0.26	0.11	0.25	0.21	0.10	0.03	0.08	0.08	0.34	0.40
FeO	6.13	7.48	7.37	6.43	6.49	6.29	6.56	5.37	6.32	6.75
MnO	0.06	0.05	0.05	0.05	0.06	0.05	0.02	0.03	0.04	0.05
MgO	4.46	4.87	4.44	4.63	4.43	4.46	4.87	4.03	4.39	4.34
CaO	8.52	9.87	8.49	8.88	8.31	8.91	9.68	8.08	8.19	8.20
Na ₂ O	9.18	8.38	9.08	8.84	9.05	9.20	8.79	9.66	9.82	9.54
K ₂ O	0.31	0.17	0.19	0.31	0.14	0.20	0.18	0.23	0.16	0.16
Total	99.73	99.75	99.67	99.81	99.34	99.88	99.70	99.31	99.63	99.72
Si	1.937	1.948	1.955	1.953	1.940	1.939	1.936	1.934	1.943	1.952
Ti	0.028	0.030	0.025	0.029	0.028	0.031	0.035	0.038	0.026	0.025
Al	0.656	0.581	0.613	0.623	0.659	0.645	0.609	0.685	0.643	0.629
Cr	0.007	0.003	0.007	0.006	0.003	0.001	0.002	0.002	0.010	0.011
Fe	0.183	0.226	0.222	0.192	0.195	0.188	0.197	0.160	0.189	0.202
Mn	0.002	0.002	0.002	0.002	0.002	0.002	0.001	0.001	0.001	0.002
Mg	0.238	0.262	0.238	0.247	0.237	0.238	0.261	0.215	0.235	0.232
Ca	0.326	0.381	0.327	0.340	0.319	0.341	0.373	0.309	0.315	0.315
Na	0.636	0.586	0.633	0.613	0.629	0.637	0.613	0.669	0.682	0.663
K	0.014	0.008	0.009	0.014	0.006	0.009	0.008	0.010	0.007	0.007
Total	4.028	4.026	4.030	4.018	4.018	4.030	4.034	4.024	4.050	4.038
100Fe#	43.54	46.28	48.22	43.79	45.11	44.17	43.04	42.78	44.68	46.60

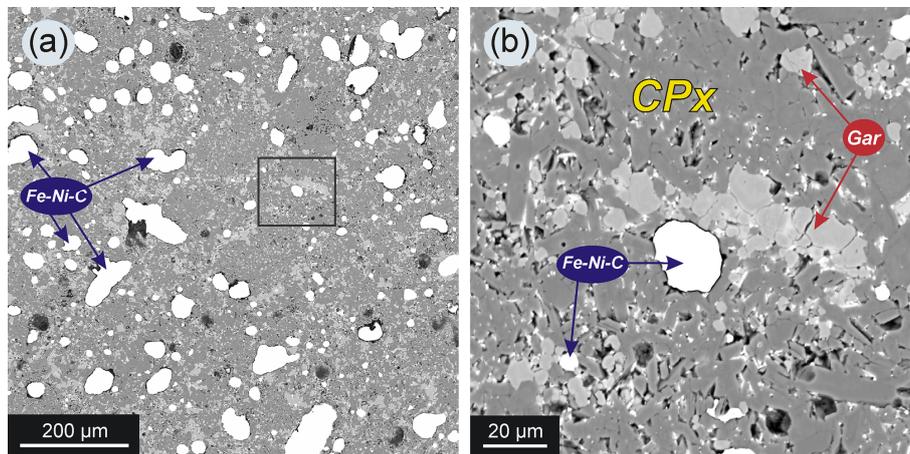


Figure 6. SEM images of synthetic silicates from the upper part of the chamber (assembly configuration as in Fig. 1f) from experiment 2–47 with 15 wt % basalt addition to the system: (a) general view, Fe–Ni–C – quenched metal (bright spots); (b) magnified view of area shown with a black square in (a) – silicates are represented by garnet (Gar) and clinopyroxene (Cpx). Note that metal drops in the silicate matrix are not interconnected.

periment, we need to understand if these minerals reflect the same conditions under which the diamond grew.

The process of formation of silicate minerals in our experiments is as follows. At the beginning of the experiments, basalt particles were present in the form of a uniformly distributed mass of fine-grained silicate substance in the volume of the metal mixture. At the high P – T conditions of

the experiments, the basaltic material recrystallized to form garnet and clinopyroxene. Due to the significant difference in density compared with the metal melt, the silicates gradually floated upward in the crystallization chamber. In this case, the aggregation of individual particles occurred, which then filled the upper part of the reaction chamber. The size of the individual grains of synthetic silicate minerals in ex-

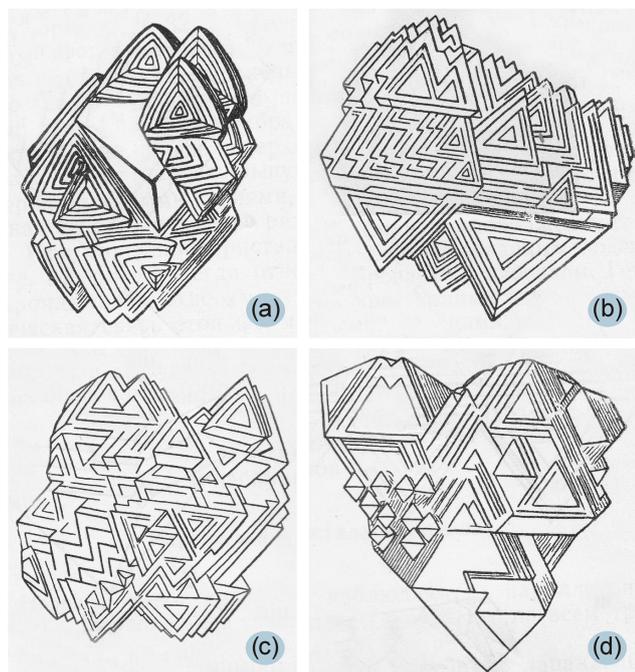


Figure 7. Parallel-growth aggregates of natural diamond crystals, original drawing from the book *Morphology of diamond* by Yu. L. Orlov (Orlov, 1963).

periments usually did not exceed $20\ \mu\text{m}$ (Fig. 6). However, these grains are larger than the original basalt particles of the starting powder ($5\ \mu\text{m}$). Often, the silicate aggregates present in the metal melt were captured by diamond crystals during growth, forming inclusions in diamond. The captured silicates can be considered as “synchronous” minerals with respect to diamond, possibly protogenetic, which is in good agreement with recent research (Nestola et al., 2017, 2019). Considering the specific appearance of the inclusion of silicates in our experiments, they can be easily distinguished by visual observation from the metal-catalyst inclusions that are dark colored and usually faceted (Fig. 2b) (Shigley et al., 1992, 1993).

With a small content of basalt in the initial mixture, silicates were present in the growth system rather as a mechanical impurity, affecting the changing morphology of the growing diamond, and the silicate phases themselves did not form representative samples. With a high content of silicate minerals in the system (more than 10 wt %), the latter formed separate volumes of silicate substance, immiscible with the metal melt, which often blocked the growth of certain areas of diamond crystals. It is important to note that in this work we use wt %, and, due to the significant difference in density between metal and silicates (more than 2 times), the volume of the silicate substance in the experiments was very significant. For example, in experiment 2–44 (Table 1) with 10 wt % basalt, the amount of silicate minerals was about 25 vol %, with respectively 75 vol % of metal. However, di-

among under such conditions grew on the seed crystal at the bottom of the chamber, most likely because during the experiment these silicate aggregations floated upwards. At the upper sections of the chamber the silicates formed a solid layer with only a little metal present as drops with sizes from less than 5 to $200\ \mu\text{m}$. We extracted the representative samples of silicates precisely from such zones (Fig. 6). In these areas, obviously, the diamond could not grow. Diamond continued to grow where the metal melt was available around the crystal and where carbon was transported to the surface of the crystal.

Thus, the micro- and macromorphology of diamonds synthesized in our experiments (with silicate phases in the growth medium) largely reflect the model situation created under the experimental conditions in a very limited amount of metal melt. Under natural conditions, there is a wide variety of options. In our experiments, diamond growth occurred on the seed crystal fixed in the cell material – i.e., silicate phases moving in a metal melt could easily block its growth. In natural conditions it is more likely that diamond crystallization centers occur in the volume of the metal melt.

4.3 Application to natural geological environments

In order to transfer the results of our experiments to natural objects, we should refer to the latest data on the findings of metallic inclusions in large crystals of natural sublithospheric diamonds (Smith et al., 2016). The diamond crystallization process, according to the proposed hypothesis, is based on the formation and evolution of metal pockets which are formed during subduction. Such pockets are nothing but volumes of metal melt. In a first approximation, our high-pressure experimental cells also represent small volumes of metal melt in which carbon is present and diamond crystallization occurs. In this case, an important role could also be played by the fluid phase, which is always present in metal–carbon melts and is represented mainly by hydrocarbons (Zhimulev et al., 2015). Recent papers by Smith et al. (2016, 2018) discuss the presence of a fluid phase (CH_4) in inclusions in diamonds and consider diamond growth in the presence of metallic melt.

In our experiments, we artificially added silicate components to the system, simulating the immiscibility of the metal melt and silicate substance, and the silicate additives were small in size and evenly distributed in the volume of the metal melt. We can expect that in natural conditions, during the formation of metal pockets, the silicate substance will be mixed with a metal melt at certain times. An example would be the finding of polyphase inclusions with the simultaneous presence of silicate, carbonate, and sulfide components (Mikhno and Korsakov, 2015). At the same time, small particles of silicate minerals can behave in a similar way as in the experiment: the process of consolidation of individual silicate grains into larger aggregates will occur very quickly, while they will float upward in the metal melt. As a result, after a

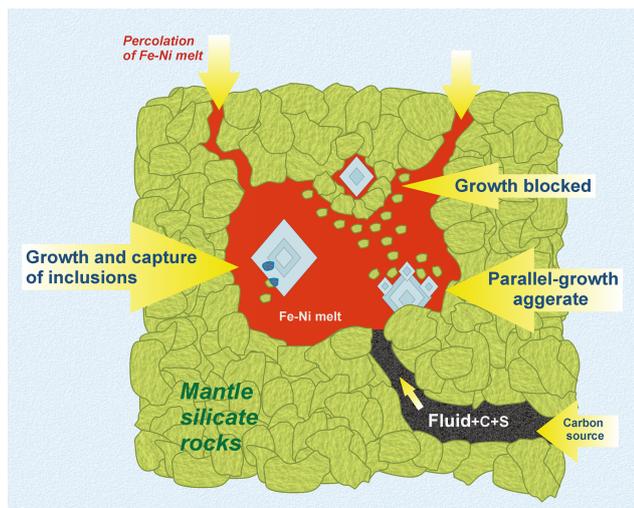


Figure 8. Scheme illustrating the growth of diamond crystals in metal melt pocket. Percolation of the Fe–Ni metal through silicate rock in the mantle (shown with arrows in the upper part) forms the metal pocket. Fluid could act as a transporter of carbon and other light elements, like sulfur, to the pocket. High-quality large diamond crystals can grow in the volume of the metal melt that is almost silicate-free (large crystal on the left); rare inclusions of silicate minerals, which are present together with diamond in the metal melt, can be captured by the crystal. Parallel-growth aggregates of diamond can form when there are many silicate phases in the metal around the growing diamond (crystal at the bottom). Diamond growth is completely interrupted by silicate minerals floating in the metal melt – such case is most likely to occur in the upper part of the pocket (small crystal at the top).

fairly short period of time, the metal pocket became silicate-free. In such conditions and with the presence of a carbon source, growth of high-quality diamond crystals which are almost free from inclusions of silicate minerals could occur. In certain areas of the pocket where diamond crystals are surrounded by a large amount of a silicate substance, diamond crystallization becomes much more complicated. It is important to emphasize that since chemically solid silicates interact neither with diamond nor with metal melt, the mechanical interaction of diamond and silicates will inevitably occur, resulting in the capture of silicate minerals in the form of inclusions, then the complication of the morphology of the growing facets of diamond, or such phenomena as the formation of parallel-growth aggregates (Fig. 8).

It seems quite possible that the size of the metal pockets in the Earth’s mantle can be very large: judging by the findings of sublithospheric diamonds weighing several hundreds and thousands of carats (Scarratt and Shor, 2006; Smith et al., 2016), which are often only fragments of larger crystals, the size of evolving drops of metal melt in the Earth’s mantle may reach gigantic sizes. In this case, the crystallization centers, i.e., diamond seed crystals, could occur in the volume of the metal, thus allowing for the growth of a diamond crystal

in all directions. At the same time, the growth rates of natural diamonds can be much lower compared with model experiments. All this, ultimately, will lead to the formation of high-quality octahedral diamond crystals with a small amount of inclusions. Changes in the chemical composition of the melt occurring in natural environments, carbon delay, or other factors will obviously affect the dynamics of diamond crystal growth in the metal melt, up to the interruption of growth and even dissolution of crystals – such processes are reflected in a complex zonation which is observed in almost every natural diamond crystal (Beskrovanov, 2000; Kaminsky and Khachatryan, 2004; Nasdala et al., 2005; Shirey and Shigley, 2013; Bulanova et al., 2017). Regarding the issue of stability and dissolution of diamond crystals, it is worth noting that the presence of silicate substances in a heterogeneous metal–silicate system can also contribute to the formation of complex diamond dissolution surfaces, which has been demonstrated experimentally (Chepurov et al., 2018b).

It is important to keep in mind that metal melts are sufficiently mobile liquids that can easily migrate under certain conditions (Fig. 8). It is believed that Fe melt can percolate through the silicate matrix if its amount exceeds 5 vol % (Yoshino et al., 2004). The critical amount of the metal melt fraction that forms a network is 6 vol % for a dihedral angle of 60–85°. At the same time, it is estimated that the residual fractions of the metal melt (metal pockets) can reach 2–5 vol % (Walte et al., 2011). There is another mechanism that contributes to the percolation of metal melts through the interstices in a silicate rock, and we have recently studied this effect experimentally (Zhimulev et al., 2018). This process will stimulate the evolution of metal pockets in the Earth’s mantle. At the same time, the migrating metal melt is saturated with carbon and other light components, such as sulfur, forming a complex Fe–Ni–C–S liquid in which diamond crystals can grow (Zhimulev et al., 2012). This fact is very interesting for mantle geology, because it demonstrates the growth of diamond from metal–sulfide melts. It is worth noting that sulfide inclusions are one of the most common in diamonds, especially in central areas of the crystals (Harris and Gurney, 1979; Bulanova et al., 1990).

Presenting the models of diamond crystallization and its capture of inclusions of silicates, oxides, sulfides, and metals during growth, one should take into account the possibility of transformation of these phases in postcrystallization conditions. So, after the completion of crystallization, diamonds for hundreds of millions and even billions of years were stored in the Earth’s mantle. During this time, inclusions of silicates, oxides, and sulfides remained completely inert “prisoners” of diamond, which is confirmed by the experimental data (Fedorov et al., 2006). On the other hand, inclusions of metals at mantle P – T conditions are able to move through the volume of a diamond crystal by recrystallization of the diamond substance, thereby reaching the surface of the crystal and leaving the diamond (Chepurov et al., 2000; Fedorov et al., 2005). Thus, this process of self-purification of

diamond from inclusions of metal melts can serve as one of the explanations for the relatively rare findings of native iron inclusions in natural diamonds.

The mechanisms of crystallization of diamond and silicates discussed above are relevant in the general context of the Earth's evolution. It is known that most inclusions in natural lithospheric diamonds are of an ancient age (Gurney et al., 2010), and peridotitic inclusions reach an age of more than 3 Ga (Richardson et al., 1984). During this period, the continental mantle was reduced, corresponding to the equilibrium of silicates with the Fe–Ni phase (Kadik et al., 2004), and for peridotite xenoliths the oxygen fugacity level is estimated to be close to the Fe–Ni melt stability line (Shirey et al., 2013). It is important to note that metal–carbon melts can be stable under more oxidized conditions than the corresponding solid metal phase; thus, diamonds in the Fe–Ni–C melt can crystallize even with wüstite (Fedorov et al., 2002). Therefore, if diamonds with inclusions of Fe–Ni and Fe–Ni–sulfides are considered as evidence of diamond growth with the participation of metal liquids under upper-mantle conditions, it can be assumed that the formation of such diamonds, as well as crystallization of coexisting silicate minerals, could occur precisely under the highly reduced conditions in the ancient mantle of the Earth where iron and nickel existed in the form of metal or metal–sulfide melts.

Silicate minerals in our experiments were formed in highly reduced conditions and consisted mainly of garnet and clinopyroxene. Clinopyroxene is characterized by a high Na content and garnet is a low-chromium variety with a high Ca content. Therefore, the chemical composition of the synthetic minerals is close to the typical inclusions of the eclogitic assemblage in natural diamonds, namely to omphacite clinopyroxene and grossular–pyrope–almandine garnet (Fig. 9). Eclogites are considered derivatives of subducting basalts, and, consequently, they pass a long and complex way deep into the planet. The role of subduction processes in the genesis of diamond is still insufficiently studied and little discussed in the literature, especially from the standpoint of understanding the role of Fe–Ni and sulfide melts. It is known that Fe–Ni sulfides (and even native iron) are usually present in diamond-bearing eclogites where these sulfide inclusions are more often present in the form of intermittent rims around diamond (Bulanova et al., 1990). The results of our experiments lead us to consider the possibility of crystallization of some diamonds during subduction at lithospheric-mantle conditions with the participation of metallic liquids. Confirmation of this can be the finding of inclusions of metals in diamonds, which cannot be classified as sublithospheric ones according to the whole set of features.

In later epochs, as planet Earth evolved, highly reduced conditions prevailed only in the deeper parts of the Earth. It is assumed that the formation of the metal phase begins at depths greater than 250 km (Frost and McCammon, 2008; Rohrbach et al., 2007, 2011). Smith et al. (2016) proposed a mechanism of diamond formation with the participation

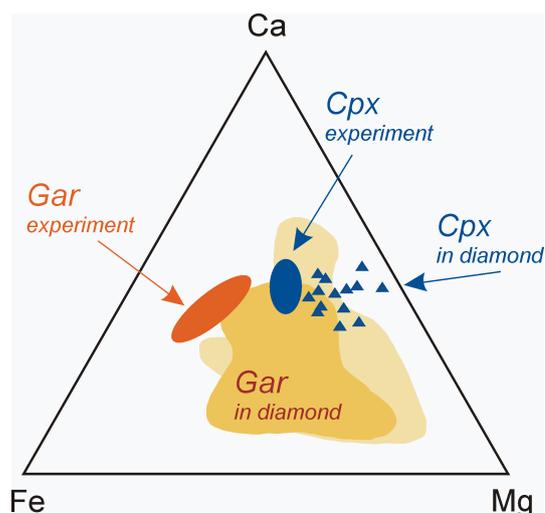


Figure 9. Chemical composition of synthetic garnet (oval-shape field “Gar experiment”) and clinopyroxene (oval-shape field “Cpx experiment”) of this study compared to silicate inclusions of eclogitic paragenesis in natural diamonds: grossular–pyrope–almandine garnet (“Gar in diamond”) and omphacite clinopyroxene (triangles). Data on natural samples are from Sobolev (1977), Harris and Gurney (1979), Davis et al. (2003), and Liu et al. (2009).

of metal liquids, suggesting the possibility of subduction of crustal material to the depths of the lower mantle where native metals and, accordingly, silicate minerals of higher pressures become stable. Highly reduced conditions at super depths of the Earth's mantle stimulate the crystallization of diamond in metal melts, which are stable under these conditions and are a favorable environment for the growth of high-quality and large diamond crystals. At the same time, diamond as an inclusion will capture silicate minerals that are typical for these depths. This process is probably happening now.

5 Conclusions

This experimental study on the joint crystallization of diamond in the presence of silicate minerals has shown that the presence of silicates in the crystallization medium does not reduce the chemical ability of metal melts to catalyze the conversion of graphite into diamond, and, morphologically, growth of diamond on seed crystals occurred mainly in the form of octahedra. A more difficult task was to grow high-quality diamond single crystals on the seeds in a heterogeneous system. The main effect of silicates on diamond growth in the metal–carbon system was the loss of morphological stability of the {111} faces, which starts at 5 wt % of silicate substance in the crystallization medium. At a very high content in the system, the silicate phases completely block some areas of the crystallization chamber, preventing the penetration of the metal melt into them. In this case, the

growth of a diamond crystal on the seeds stopped. Thus, the diamond growth in a complex heterogeneous environment, consisting of metal–carbon melt and silicate components, depends not only on temperature, pressure, and chemical composition of the system but largely on the transport capacity of the medium.

The present paper shows the process of crystallization of diamond in a metal melt together with garnet and clinopyroxene, whose chemical composition is similar to minerals of eclogites. This means that, under natural conditions, diamonds could grow in the pockets of metal melt surrounded by such silicate rock. In this case, a key factor for the progress of diamond growth will be the maintenance of highly reduced conditions in which the metal melt environment for crystallization of diamond exists. The described mechanism can occur not only in modern conditions in the deepest mantle but probably could have taken place in the ancient continental mantle of the Earth, which was reduced, corresponding to the equilibrium of silicates with a Fe–Ni melt.

Data availability. No data sets were used in this article.

Author contributions. AIC and VS designed and carried out the experiments. EZ performed microscopic studies. AIC, AnC, and J-MD prepared the paper with contributions from all coauthors.

Competing interests. The authors declare that they have no conflict of interest.

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